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# Graphing periodic trends lab answer key

Thank you for your participation! Thank you for your participation! Replicated activity: Purpose of graphing trends: Use the distance table to identify and explain the timely trends, including nuclear and ionic reaction, electrical negativeness and tain energy. Discussion: There are many periodic trends that you find as you go across a period or a group/family distance table. Graphing will help to explain these trends to different elemental characteristics. Reference: Anstrom-a unit of 1 meter of distance equal to 10-billionth. Nuclear-response-view-distance from the Center for Electrical Energy Levels. Ionic-response-to-the-center-to-electric energy levels. Tain Energy-Energy needed to remove an electrical from an atom. Electrical negative-energy with which an atom can attract data: The first 18 elements feature nuclear Ionic-Tain nuclear electro-negative element Radas Radas Energy Number (Palong) (Engsteramus) (Engsteramus) (KJ) H 1 0.79 0.012 314 567 ؟ Lee 3 2.05 0.76 124 0.98 Ho 4 1.4 0.35 215 1.57 B 5 1.17 0.23 191 2.04 C 6 0.91 0.16 2 60 2.55 N 7 0.75 0.13 335 3.04 O 8 0.65 1.4 314 3.44 F 9 0.57 1.33 402 4 Ne 10 . Na 11 2.23 1.02 119 0.93 mg 12 1.72 0.72 176 1.31 Al 13 1.82 0.535 138 1.61 Si 14 14 1.46 0.4 1 88 1.9 s 15 1.23 0.38 242 2.19 S 16 1.09 0.37 239 2.58, 17 0.97 1.81 299 3.16 is 18 . Using the following statistics, build the following line graphs: (give each graph a title and label both xaxis and y axis. Contact your points. Predicts plot points for elements with question marks. Activity 1: Graphing Trends 1. Nuclear Response (y) vs. Nuclear Number (x). 2. Ionic Radas (y) vs. Nuclear Number (x). 3. Tain Eenergy (y) vs. Nuclear Number (x). 4. Electrical negative (y) vs. nuclear number (x). Activity 2: Graphing is a modified distance table for the elements 1-18 below the distance trends. Each element reflects the atomic size using the data provided above for its close-up model for each element. Try to take as much relative size as possible into the account and the electronic cloud. Bohr Model Nuclear Size 1 2 3 4 5 6 8 9 10 11 12 13 14 15 16 17 18 Use your graphs to answer the following questions. Apply the correct answer and explain why it is correct. 1. The response [added, decreased] as you go across (L to R) a period? Explain why-2. When you go down to a family, there is a reason for the decline . Explain why-3. Electricity usually needs to remove an electrical from an atom (added, decreased) as you go across a period? Explain why this happens. 4. Why is the electro-negative of hydrogen smaller than the electro-negative? Activity 3: Graphing Distance Trends 5. Atom circuit with the largest atomic reaction (size) in each group: a. b. d. aluminum, sulphur, phosphorus, zererich, bismuth, hydrogen iron, lytium, Berem, Baralim, Bromana. The smallest ionic reaction in each group will be that atom circuit: a. b. d. magnesium, chlorine, silicon latim, seisium, potassium fluorine, iodine, chlorine calcium, bromana, cobalt 7. Circuit this atom which will need to remove an EA of maximum energy. b. c. latim, potassium, robadyum sodium, chlorine, silicon polonium, oxygen, sulfur fluorine, iodine, chlorine, the correct response in the bracket below the chlorine (e) circuit. 8. A period from left to right is filled: [p , N, e-] increase, thus [p , N, e-] to the center of the atom and [increasing, decreased] nuclear reaction. Because of this increasing the electronics, the strong force are to nuclear [recover, lose] electrons as you go across the distant table. 9. ایک دھاتی خاندان نیچ□ جا: [p + ن e-] □چھوٹ, □بڑ] [□ردا س بنانا□, جو□ری میں اضافہ□ [توانائی کی سطح کی تعداد]. Because electrons are far from, they tend to get more easily. So metals are [more, less] chemically active as your family go down. 10. Going down to a non-metallic family: This institution gets [the space, away] from the external energy level, thus [increasing/reducing] the metal-free electrons pull on. This non-metals [more, less] chemical is active because you go down to a family. 4 4

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